**SAMPLE CHAPTER**

**Chapter 3**

**Statistical Modeling for Adsorption of Congo Red onto Modified Bentonite**

***Mohan Rao1\*, Karla Jones1, Vivian Vance2***

*1Dept. of Chemical Engineering, Bapatla Engineering College, Bapatla, India*

*2 Faculty of Technology, University College of Technology, Hyderabad, India*

*\*Corresponding author. E-mail: marao@gmail.com*

**ABSTRACT**

Estimation of total content of heavy metal in different horizons of acid sod-podzolic soils of Kirov region (Russia) covered with natural vegetation (forest, meadow) and used as arable land is spent by a method of atomic-absorption spectroscopy. The content of total cadmium in differ- ent horizons of sod-podzolic and podzolic loamy soils fluctuated from 0.66 to 1.11 mg/kg. The content of exchangeable cadmium in soils fluctuated from 0.01 to 0.30 mg/kg in different horizons. The greatest content of exchangeable Cd is noted in accumulating eluvial parts of a profile and in mother rock, the least in illuvial horizons. In humus horizons of meadow and fallow land soils, its content is statistically higher than in arable land soils (0.17±0.02 and 0.08±0.02 mg/kg, accordingly). Ratio of exchangeable cadmium in its total content varied largely. In the top part of a profile, it fluctuated from 10 to 40%, in illuvial parts drop down to 1–11% and increased again at approach to mother rock. During the growth season, content of exchangeable Cd in the top horizons of forest podzolic soil can vary by 4–5 times, sometimes reaching the critical values close to maximum allowable concentration (MAC) (0.2 mg/kg). In fallow land sod-podzolic soil the maximum is revealed in the middle of a season (1.0 mg/kg) that is almost 10 times higher than in autumn. On an arable land under clover crops, total cadmium content is minimum in the beginning and the middle of a season, and sharply rises in the autumn to the values close to MAC (0.8 mg/kg). It has not revealed any significant correlations between content of cadmium with other agrochemical properties of arable land soils. In covering loams of different districts of the Kirov region, the amount of total Pb fluctuated from 5.0 to 43.0 mg/kg.

**3.1 INTRODUCTION**

Synthetic dye materials from various industries, such as textile, leather, plastic and paper, pollute bodies of water and act as an eco-toxic hazard that contributes to serious environmental problems. Azo dyes, like Congo red (CR), have been known to cause allergic reactions and various toxic effects on human and aquatics [1-5] and must be removed before they are discharged into water streams.

Dyes have a complex structure and are resistant to various degradation methods that have posed technological challenges for many decades. Compared to several conventional methods of separation, like oxidation, membrane separation, precipitation, coagulation and electro dialysis for color de contamination, adsorption is proven to be most popular process due to its simplicity, high efficiency, and easy recoverability and reusability of the adsorbent, [6, 7]. Numerous adsorbents have been developed so far, and activated carbon is most popular among all, due to its high adsorption capacity. But it is expensive and not economical to regenerate [8, 9]. This has led to a search for alternative materials that are cost effective and can replace activated carbon. In recent years, extensive research has been carried out for low-cost adsorbents. Various naturally occurring materials and industrial waste materials were tested. Among them, bentonite has been proven to be a promising material for the adsorption process in its pure and modified forms [9-11].

In this work, raw bentonite was modified by interacting with an acid solution to synthesize a new absorbent (modified nentonite) to remove Congo red dye from the aqueous solution. The objective of the present work is to develop a successful model for CR removal of modified Bentonite with the help of statistical tools.

**3.2 EXPERIMENTAL METHODS AND MATERIALS**

**3.2.1 Adsorbent**

The bentonite was activated by adding concentrated H2SO4 (1:1 w/v) with constant stirring. The material was kept in a hot air oven at 110°C for 12 hours. This material was washed with distilled water and was soaked in 2% NaHCO3 solution overnight to remove the residual acid. Then the material was washed with distilled water until the pH of the adsorbent reached slightly above 7. Finally, it was dried in a hot air oven at 1100C for 4 hours. The particle size was determined by sieving the dried material at 125µm. The sieved adsorbent was stored in an airtight container for further experimentation.

**3.2.2 Adsorbate**

Congo red is the first synthetic azodye produced that is capable of dyeing cotton directly. A stock solution of the dye with a concentration of 1000 mg L-1 was prepared with Millipore water, and it is diluted to achieve the working solutions. The adsorbent dosage was measured accurately with an analytical balance (SHIMADZU – AX200). The pH of the solution is measured with a digital pH meter (ELICO-L1 612) and varied by using 0.1 N HC1 and 0.1 N NaOH solutions. The solution with added adsorbent is agitated with a Remimake Temperature Controlled Orbital Shaker (REMI - CIS 24 BL). At the end, the samples were collected and centrifuged to remove the suspended solid particles using REMI C 24 centrifuge. The clear liquid was collected and analyzed with UV-Visible Spectrophotometer (SYSTRONICS-117) at a wavelength of 498 nm.

All determinations were performed in triplicate per experiment. The amount of CR adsorbed onto modified bentonite was measured in terms of uptake, qt (mg of CR adsorbed per one gram of adsorbent) and percent removal %R which are calculated using the following equations:

 (1)

 (2)

**3.3 RESULTS AND DISCUSSION**

**3.3.1 Effect of Operating Parameters**

***3.3.1.1 Effect of Amount of Adsorbent***

The amount of dye uptake prominently depends on the surface area available. Experiments were conducted to examine the extent of its dependence and fix the optimum dosage. Different amounts of adsorbent were added to 50 ml of 100 mg/L solution at 30°C, and the results are displayed in Figure 3.1. The dye removal percentage increased with an increasing amount of adsorbent from 0.02 to 0.1 g and slowed down up to 0.2 g. The further increase in the amount of adsorbent did not affect the removal significantly.

**FIGURE 3.1** The effect of adsorbent dosage on the color removal. (*Source*: Reprinted with permission from Ref. [22]. © 2012 Elsevier.)

***3.3.1.2 Effect of Contact Time***

Experiments were conducted by adding 0.1 g adsorbent to 50 ml of 100 mg/L solution and agitated at fixed temperature. Samples withdrawn at different time intervals and were analyzed. This was repeated at temperatures (20, 40 & 500C) and the results were shown in Figure 3.2.

**FIGURE 3.2** The effect of solution temperature on adsorption.  (*Source*: Reprinted from Ref. [10]. [http://creativecommons.org/licenses/by/4.0](https://creativecommons.org/licenses/by/4.0/))

Adsorption was rapid in the beginning up to first 50 min and thereafter it is stabilized.Finally, it attained equilibrium at 180 min. There is no considerable change in % adsorption at different temperatures, but it is slightly less at 50°C (96%) compared to other temperatures (98%).

***3.3.1.3 Effect of pH***

Effect of pH in the range of 2 - 12 was examined with 100 mg/L solution. The equilibrium dye uptake at these conditions are determined and tabulated (Table 3.1). There is a slight decrease in uptakewithpH. The insignificant change may be due to retention of H+ or OH**-**of clay fraction made the solution tend to neutral. Similar results were observed in case of Congo red removal with modified-bentonite [7] and adsorption of Ni, Zn and Pb onto bentonite [12].

**TABLE 3.1** Effect of pH on the Adsorption Capacity of Modified Bentonite

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| pH | 2 | 4 | 6 | 8 | 10 | 12 |
| uptake | 49.99 | 49.74 | 46.32 | 47.03 | 48.41 | 47.61 |

*(Source*: Reprinted with permission from Ref. [22]. © 2012 Elsevier.)

***3.3.1.4 Effect of Temperature***

The effect of temperature on adsorption of CR onto MB was presented in Figure 3.3. % removal increased with increase in solution temperature from 20oC to 50oC. This shows that the process is endothermic in nature.

The increase in percentage removal at higher temperature may be due to a greater kinetic energy acquired, resulting in an easier diffusion from the bulk solution onto the surface of MB.

**3.3.2 Statistical Modeling and Analysis**

The factors that influence the sorption process are effluent concentration (Co), dosage (Do), temperature (T), pH, time, and speed of agitation. The agitation speed is fixed at 150 rpm. Equilibrium time can beestimated from equilibrium studiesand is taken a 4hrs.The negligible influence of pH is observed in parameter studies.

The scheme of experiment in soil culture:

1. Control soil (without entering HM);

2. Control soil + 1 mg/kg of cadmium;

3. Control soil + 5 mg/kg of cadmium;

Total number of runs (Nr) required in totality are estimated as:

 (3)

Experiments were conducted according to the design matrix generatedand the measured responses were presented in Table 3.3.

**TABLE 3.2** Experimental Range and Levels of Independent Variables

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
| **Variables** | **Code** | **Symbol** | **- α** | **-1** | **0** | **1** | **+α** |
| Initial Concentration, mg/L | A | Co | 20 | 218.64 | 510 | 801.36 | 1000 |
| Dosage, g | B | Do | 0.01 | 0.109 | 0.255 | 0.401 | 0.5 |
| Temperature, oC | C | T | 20 | 26.08 | 35 | 43.92 | 50 |

**TABLE 3.3** Experimental Design Matrix

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Run** | **Concentration, mg/L** | **Dosage, g** | **Temperature, oC** | **Uptake,**  **mg/g** |
| 1 | 510.00 | 0.500 | 35.00 | 49.5185 |
| 2 | 510.00 | 0.010 | 35.00 | 424.691 |
| 3 | 801.36 | 0.109 | 43.92 | 281.609 |
| 4 | 510.00 | 0.255 | 35.00 | 85.3098 |
| 5 | 510.00 | 0.255 | 35.00 | 84.7756 |
| 6 | 20.00 | 0.255 | 35.00 | 3.36538 |
| 7 | 218.64 | 0.109 | 43.92 | 74.6764 |
| 8 | 510.00 | 0.255 | 50.00 | 84.8943 |
| 9 | 218.64 | 0.401 | 26.08 | 26.3385 |
| 10 | 510.00 | 0.255 | 35.00 | 85.292 |
| 11 | 510.00 | 0.255 | 20.00 | 77.5938 |
| 12 | 1000.00 | 0.255 | 35.00 | 159.182 |
| 13 | 801.36 | 0.109 | 26.08 | 161.098 |
| 14 | 801.36 | 0.401 | 26.08 | 88.2989 |
| 15 | 801.36 | 0.401 | 43.92 | 91.7675 |
| 16 | 510.00 | 0.255 | 35.00 | 85.7597 |
| 17 | 218.64 | 0.401 | 43.92 | 26.6857 |
| 18 | 218.64 | 0.109 | 26.08 | 63.7337 |
| 19 | 510.00 | 0.255 | 35.00 | 85.2802 |
| 20 | 510.00 | 0.255 | 35.00 | 84.8528 |

Various models were tested with the experimental data to obtain the regression equations. And the model adequacy was tested with sequential F-test, lack-of-fit test and otheradequacy measures were used for selecting the best model [19, 20].Analyzing the measured responses, the fit summary output indicated that the quadratic polynomial model was significant for the present system. General form of the quadratic model is

 (4)

Where q is CR uptake, b0 = constant coefficient,,­,, are the interaction coefficients of linear, quadratic and second order terms respectively. n is the number of factor and is the error [21,22].

The uptake of CR varied from 3.36 (minimum) to 424.69 mg/g (maximum) in the present study. Theratio of minimum to maximum uptake of congo red removal was 126.39, which was greater than 10 suggesting that transformationwas required in the present system. Since CR removalin the present investigationrepresented right skewness with non-zero positive values, a naturallog transformation was applied to the experimental data.

Sequential model sum of squares (Table 3.4), lack of fit tests (Table 3.5) and model summary statistics ( Table 3.6) were carried out to check the adequacy of the model for CR removal by the MB. p values for the regressions were lower than 0.01 (see Table 3.4) suggesting a model of quadratic. The model summary statistics (Table 3.6) indicates higher regression coefficient (R2 = 0.9284) for the quadratic model with the minimum standard deviation (0.36).

**TABLE 3.4** Sequential Model Sum of Squares

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **Source** | **Sum of Squares** | **df** | **Mean Square** | **F Value** | **p-value**  **prob>F** | **Remarks** |
| Mean vs Total  Linear vs Mean  2FI vs Linear  Quadratic vs 2FI  Cubic vs Quadratic  Residual  Total | 373.65  13.07  0.082  3.64  1.18  0.11  391.74 | 1  3  3  3  4  6  20 | 373.65  4.36  0.027  1.21  0.30  0.019  19.59 | 13.90  0.072  9.36  15.68 | 0.0001  0.9737  0.0030  0.0025 | Suggested  Suggested  Aliased |

The significance of the regression models and individual model coefficients and the lack of fit are tested using the same statistical package. The resulting ANOVAfor the reduced quadratic models summarizes the analysis of variance of each response and shows the significant model terms. Table 3.7 shows the ANOVA result for CR removal onto the MB, with a model F- value of 14.40 implying that the model is significant (at p < 0.05). In this case, A( Co), B( Do) are highly significant model terms, A2 and B2 are significant model terms, while model values greater than 0.10 indicated that the model terms were not significant.

**TABLE 3.7** Analysis of Variance Table

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **Source** | **Sum of Squares** | **Df** | **Mean Square** | **F – Value** | **p-value**  **Prob>F** | **Remarks** |
| Model  A-Co  B-Do  C-T  AB  AC  BC  A^2  B^2  C^2  Residual  Lack of Fit  Pure Error  Cor Total | 16.79  9.16  3.85  0.062  4.521 x 10-3  0.023  0.055  2.56  0.75  7.281 x 10-3  1.30  1.30  8.809 x 10-5 18.09 | 9  1  1  1  1  1  1  1  1  1  10  5  5  19 | 1.87  9.16  3.85  0.062  4.521x10-3  0.023  0.055  2.56  0.75  7.28x10-3  0.13  0.26  1.762x10-5 | 14.40  70.73  29.71  0.48  0.035  0.17  0.43  19.73  5.77  0.056  14701.97 | 0.0001  <0.0001  0.0003  0.5049  0.8555  0.6848  0.5281  0.0013  0.0372  0.8174  <0.0001 | Significant  significant |

The final mathematical model in terms of the actual factors as determined by Design- expert software is shown below:

In terms of coded factors:

 (5)

**3.4 CONCLUSION**

Bentonite was modified and used to remove Congo red from dye effluent water. It is proven to be an effective adsorbent. The equilibrium time required is determined as 4 hours. Univariative parametric studies were conducted, and an ineffective variable (pH) was eliminated. Dye uptake increased with temperature, indicating endothermic nature of the process. The effects of various parameters and their interactions were studied using response surface methodology. Effluent concentration and dosage were identified as the most influential parameters, and the second order effects of the same are also influencing the process. The statistical quadratic model was generated, and its robustness was tested as per the standards given in the literature.

**KEYWORDS**

congo red

contour plots

modified bentonite

rating scale model (RSM)

**References**

1. Wang, L.; Li, J.; Wang, Y.; Zhao, L.; Jiang, Q.; Adsorption capability for Congo red on nanocrystal HneMFe204 (M=Mn, Fe, Co, Ni) spinel ferrites. *Chemical Engineering Journal* **2010**, doi: 10.1016/j.cej.-2011.10.088
2. Somasekhara Reddy, M. C.; Sivaramakrishna, L.; VaradaReddy, A.; The Use of an Agricultural Waste Material,Jujuba Seeds for the Removal of Anionic Dye(Congo Red) from Aqueous Medium. *J Hazard Mater* **2012**; 203:118-127.
3. Sabnis, R.W.; Handbook of Biological Dyes and Stains. Synthesis and Industrial Applications (John Wiley & Sons, New Jersey, Canada, Springer, **2010**, pp. 106-107).
4. Han, R.; Ding, D.; Xu, Y.; Zou, W.; Wang, Y.; Li, Y.; Zou, L.; Use of rice husk for the adsorption of congo red from aqueous solution in column mode. *Bioresource Technol.* **2008,** *99*, 2938-2946.
5. Raymundo, A.S.; Zanarotto, R.; Belisario, M.; Pereira, M.G.; Ribeiro, J.N.; Ribeiro, A.V.F.N.; Evaluation of sugar-cane bagasse as bioadsorbent in the textile wastewater treatment contaminated with carcinogenic congo red dye. *Brazilian Arch. Biology Technol.* **2010,** *53*, 931-938.
6. Hu, Q.H.; Xu, Z.P.; Qiao, S.Z.; A novel color removal adsorbent from hetero coagulation of cationic and anionic clays. *J. Colloid Interface Sci.* **2007,** *308*, 191-199.
7. Lian, L.; Guo, L.; Wang, A.; Use of CaC12 modified bentonite for removal of Congo red dye from aqueous solutions. *Desalination* **2009,** *249*, 797 801.
8. Sanghi, R.; Bhattacharya, B.; Review on decolorisation of aqueous dye solutions by low cost adsorbents. *Color. Technol.* **2002,** *118*, 256-269.
9. Kumar, A.; Kumar, S.; Kumar, S.; Gupta, D. V.; Adsorption of phenol and4- nitrophenol on granular activated carbon in basal salt medium: Equilibriumand kinetics. *J. Hazard. Mater.* **2007,** *147*, 155-166.
10. Xin, X.; Si, W.; Yao, Z.; Feng, R.; Du, B.; Yan, L.; Wei, Q.; Adsorption of benzoic acid from aqueous solution by three kinds of modified Bentonites. *J. Colloid Interface Sci.* **2011,** *359*, 499 -504. https://www.sciencedirect.com/science/article/pii/S0304389407000052?via%3Dihub
11. Erdem, B.; Ozcan, Ozcan A. S.; Adsorption and solid phase extraction of 8-hydroxyquinoline from aqueous solutions by using natural bentonite, *Applied Surface Science* **2010,** *256*, 5422 -5427
12. Ayari, F.; Srasra, E.; Trabelsi-Ayadi, M.; Characterization of bentonitic clays and their use as adsorbent. *Desalination* **2005,** *185*, 391-397.
13. Gottipati, R.; Mishra, S.; Process optimization of adsorption of Cr(VI) on activated carbons prepared fromplant precursors by a two-level full factorial design. *Chemical Engineering Journal* **2010,** *160,* 99-107
14. Cronje, K.J.; Chetty, K.; Carsky, M.; Sahu, J.N.; Meikap, B.C.; Optimization of chromium(VI) sorption potential using developed activated carbonfrom sugarcane bagasse with chemical activation by zinc chloride. *Desalination* **2011,** *275*, 276-284.
15. Myers, R.H.; Response Surface Methodology, Allyn and Bacon, New York, **1971**.
16. Box, G.E.P.; Hunter, J.S.; Multi factor experimental designs for exploring response surfaces. *Ann Math Statist* **1957,** *28*, 195 – 241.
17. Box, G.E.P.; Hunter, J.S.; The 2k p fractional factorial designs, parts I and II, *JTechnometrics* **1961,** *3*, 311-458.
18. Muthukumar, M.; Mohan, D.; Rajendran, M.; Optimization of mix proportions of mineral aggregates using Box-Behnken design of experiments. *Cem. Concr. Compos.* **2003**, *25*, 751 - 758.
19. Benyounis, K. Y.; Olabi, A. G.; Hashmi, M. S. J.; Effect of laser welding parameters on the heat input and weld-bead profile. *J. Mater. Process. Technol.* **2005,** *164*, 978 - 985.
20. Tarangini, K.; Kumar, A.; Satpathy, G. R.; Sangal, V. K.; Statistical Optimization of Process Parameters for Cr (VI) Biosorption onto Mixed Cultures of Pseudomonas aeruginosa and Bacillus subtilis. *Clean* **2009,** *37*, 319-327.
21. Box, G. E. P.; Hunter, J. S.; Multi factor experimental designs for exploring response surfaces. *Ann. Math. Statist.* **1957,** *28*, 195-241.
22. Ozer, A.; Gurbuz, G.; Calimli, A.; Korbahti, B.K. Investigation of nickel (II) biosorption on Enteromorphaprolifera: optimization using response surface analysis. *J. Hazard. Mater.* **2008,** *152*, 778-788.

Note:

This article has been truncated and was only created to show various elements.

Reference styles should be consistent with other chapters in the book.

Use either the American Chemical Association style used above .

OR use the Chicago style (from the *Chicago Manual of Style*) or the APA style (from the American Psychological Association). These will have in-text citations as such: (Kelly, 2004).